



Salts: The Role of Buffers (Lecture 24)



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- Certain salts, called **buffers**, can combine with excess hydrogen (H^+) or hydroxide (OH^-) ions.
 - Produce substances less acidic or alkaline.
 - Act like a chemical sponge to soak up excess acid or base, keep pH constant.
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- Buffers can be "used up". Once used up, no longer help regulate pH.

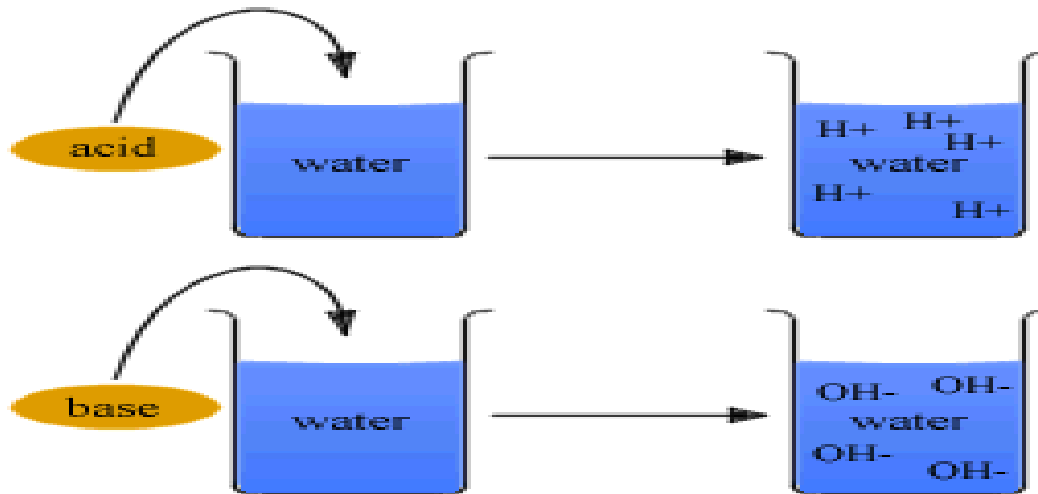


Buffers are vital to maintaining pH in organisms.

Example:

Antacids are buffers made of the salt calcium carbonate (CaCO_3).

pH: Acids & Bases and Buffers



Thank You